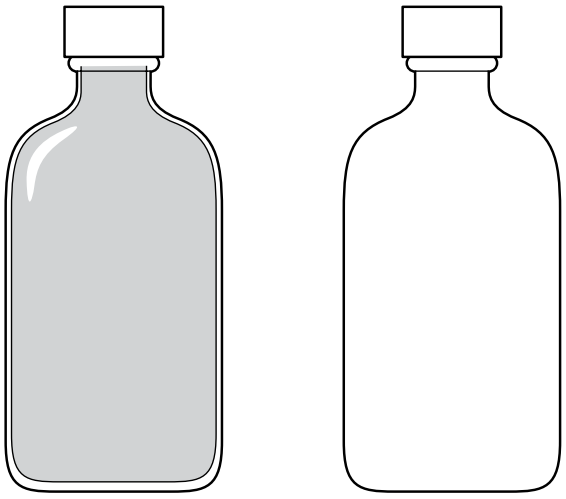
Name: ..... Form: .....

See **pages 14 and 15** of *Dissolving*

# Dissolving gases

**Gases dissolve in liquids. A hot liquid can hold much less gas than a cold one and liquids under pressure hold much more gas.**



**Q1**. The bottle of water above left has been cooled in a fridge. In the bottle above right, draw how it will look after it has warmed up for some time. Label your diagram.

**Q2.** Why was there a change when the bottle warmed up?

✎ .....

✎ .....

**Q3.** Where does oxygen enter an ocean, lake or river?

✎ .....

**Q4.** When water is boiled does it (a) make bubbles at the same rate all the time; (b) make bubbles faster as time goes on; (c) make bubbles quickly at first and then make bubbles more slowly? Explain your answer.

✎ .....

✎ .....

**Q5.** What is the gas which makes the bubbles in a fizzy drink?

✎ .....

**Q6.** A fizzy drink has a large amount of gas pushed (dissolved) into it under pressure and the pressure is maintained by the bottle top. (a) What happens to the pressure on the liquid when you open the bottle? (b) How does the pressure change affect the dissolved gas?

(a) ✎ .....

(b) ✎ .....